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| Name: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ | Date: \_\_\_\_\_\_\_\_\_ | Period: \_\_\_\_\_\_\_ |

Acid-Base Chemistry

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|  |  | 3 |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |
|  |  |  |  |  |  |  |  |  | 4 |  |  |  |  |  |  |  |  |  |  |
|  |  |  |  | 5 |  |  |  |  |  |  | 6 |  |  |  |  |  |  |  |  |
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|  |  |  | 7 |  |  |  |  |  |  |  |  |  |  |  |  | 8 |  |  |  |
|  |  |  |  |  |  | 9 |  |  |  |  |  |  |  |  |  |  |  |  |  |
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|  |  |  |  |  |  | 12 |  |  |  |  |  |  |  |  |  |  |  |  |  |
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| **Across****3.** Bases turn red indicator paper **5.** Bronsted-Lowry definition says that acids \_\_\_\_\_ H+**9.** Highest pH possible**11.** A(n) \_\_\_\_\_ solution has more hydronium than hydroxide **12.** Measured in moles per liter**13.** A(n) \_\_\_\_\_\_\_ solution has more hydroxide than hydronium **14.** Bronsted-Lowry definition says that bases \_\_\_\_\_ H+**15.** H+ | **Down****1.** Method to determine the concentration of a solution**2.** Lowest pH possible**4.** Reaction between an acid and base resulting in a salt and a water **6.** A pH of 2 is \_\_\_\_ acidic than a pH of 6**7.** The result of the acid donating its proton is the \_\_\_\_ base**8.** OH- **10.** Acids turn blue indicator paper  |