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Acid and Bases

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| **Across****4.** A solution that is neither acidic nor alkaline, such as pure water.**7.** A numeric scale used to specify the acidity or basicity of an aqueous solution.**8.**  Any substance that gives a visible sign, usually by a colour change, of the presence or absence of a threshold concentration of a chemical species, such as an acid or an alkali in a solution. **9.** A chemical compound that neutralizes or effervesces with acids and turns litmus blue; typically.**10.** Are substances that, in aqueous solution, are slippery to the touch, taste bitter, change the color of indicators. **11.** Any chemical compound formed from the reaction of an acid with a base, with all or part of the hydrogen of the acid replaced by a metal or other cation.**12.** The monovalent anion OH− consisting of one atom of hydrogen and one of oxygen. | **Down****1.** The ion H3O+, consisting of a protonated water molecule and present in all aqueous acids.**2.** Is a chemical reaction in which an acid and a base react quantitatively with each other.**3.** A solution that resists changes in pH when acid or alkali is added to it. Buffers typically involve a weak acid or alkali together with one of its salts.**5.** A molecule or other entity that can donate a proton or accept an electron pair in reactions.**6.** A measure of acidity or alkalinity of water soluble substances. |