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| Name: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ | Date: \_\_\_\_\_\_\_\_\_\_\_\_\_\_ |

Acid and Bases

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|  |  |  |  |  | 4  N | E | U | T | R | A | L |  |  | U |  |  | F |  |  |
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| **Across**  **4.** A solution that is neither acidic nor alkaline, such as pure water.  **7.** A numeric scale used to specify the acidity or basicity of an aqueous solution.  **8.**  Any substance that gives a visible sign, usually by a colour change, of the presence or absence of a threshold concentration of a chemical species, such as an acid or an alkali in a solution.  **9.** A chemical compound that neutralizes or effervesces with acids and turns litmus blue; typically.  **10.** Are substances that, in aqueous solution, are slippery to the touch, taste bitter, change the color of indicators.  **11.** Any chemical compound formed from the reaction of an acid with a base, with all or part of the hydrogen of the acid replaced by a metal or other cation.  **12.** The monovalent anion OH− consisting of one atom of hydrogen and one of oxygen. | **Down**  **1.** The ion H3O+, consisting of a protonated water molecule and present in all aqueous acids.  **2.** Is a chemical reaction in which an acid and a base react quantitatively with each other.  **3.** A solution that resists changes in pH when acid or alkali is added to it. Buffers typically involve a weak acid or alkali together with one of its salts.  **5.** A molecule or other entity that can donate a proton or accept an electron pair in reactions.  **6.** A measure of acidity or alkalinity of water soluble substances. |