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Acids And Bases

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| **Across**  **2.** Any solution in which the hydrogen-ion concetration is greater then the hydroxide ion concentration.  **5.** A basic solution  **10.**  number used to tell whether a substance is an acid, a base, or if the substance is neutral.  **11.** Containing or having the properties of a \_\_\_\_ or having a pH of more than 7.  **12.** The point in a titration at which the indicator change color.  **13.** a base that completeely dissociates into metal ions and hydroxide ions in an aqueous solution.  **15.** a substance that is completely (almost completely) ionized in an aqueous solution.  **16.** A substance that is slightly ionized in an aqueous solution  **17.** Any acid that contains three ionizable protons. (hydrogen ions)  **18.** A substance that can donate a pair of electrons to form a covalent bond.  **20.** A reaction in which an acid and base react in an aqueous solution to produce salt and water. | **Down**  **1.** The point in titration where the number of moles of hydrogen ions equals the number of moles of hydrogen ions.  **3.** Any solution in which the hydroxide ion concentration is greater then the hydrogen ion concentration.  **4.** The particle formed when a base gains a hydrogen ion.  **6.** any acid that contains two ionizable protons. (hydrogen Ions)  **7.** An aqueous solution in which the concentrations of hydrogen and hydroxide ions are equal: has a pH of 7.0  **8.** Any acid that can accept a pair of electrons to form a covalent bond.  **9.** Any acid that contains one proton. (hydrogen ion)  **14.** Containing or having the properties of an \_\_\_\_ or having a pH of less than 7.  **19.** A base that reacts with water to form the hydroxide ion and the conjugate acid of the base. |