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Acids & Bases

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| **Across****4.** Have pH = 7**6.** pH = -log[H+]**8.** Acids that ionize completely in solution.**9.** Chemicals that change color in the presence of acids or bases.**11.** Acids that only ionize partially in solution.**15.** The species produced when a base accepts a hydrogen ion to form an acid.**16.** pOH = -log[OH-]**18.** Low pOH and high pH**19.** An indicator that is used to determine if a solution is acidic or basic. Red litmus turns blue for bases, while blue litmus turns red for acids.**20.** The species produced when an acid donates a hydrogen ion to form a base.**21.** Acid contains H and dissociates to produce H+ ions in aqueous solution, while a base contains OH and dissociates to produce OH- ions in aqueous solution.**22.** Low pH and high pOH | **Down****1.** OH-**2.** When acids and bases ionize - fall apart - in solution to form electrolyte solutions.**3.** H3O+ (can be used interchangeably with H+)**5.** A measure of the strength of an acid or base solution which is based on the amount of H+ ion.**7.** Have pH < 7**10.** Have pH > 7**12.** H+**13.** Bases that ionize only partially in dilute aqueous solution to form the conjugate acid and hydroxide ions.**14.** A measure of the strength of an acid or base solution which is based on the amount of OH- ion.**17.** Bases that dissociate entirely into metal ions and hydroxide (OH-) ions in aqueous solution (Arrhenius base). |