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CHEMICAL REACTIONS VOCABULARY

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| **Across**  **2.** are diagrams that show the bonding between atoms of a molecule and the lone pairs of electrons that may exist in the molecule.  **5.** a substance that takes part in and undergoes change during a reaction.  **12.** a strong force of attraction holding atoms together in a molecule or crystal, resulting from the sharing or transfer of electrons.  **14.**  an atom or a molecule in which the total number of electrons is not equal to the total number of protons  **15.** when an electron relocates from an atom to another such chemical entity.  **16.** are a type of chemical bond where two atoms share a pair of electrons with each other. | **Down**  **1.** the minimum quantity of energy that the reacting species must possess in order to undergo a specified reaction.  **3.** process describes a process or reaction in which the system absorbs energy from its surroundings; usually, but not always, in the form of heat.  **4.** a chemical reaction that releases energy by light or heat  **6.** a mathematical relationship or rule expressed in symbols.  **7.** a chemical bond that involves the sharing of electron pairs between atoms.  **8.** the state or power of being reactive or the degree to which a thing is reactive.  **9.** a type of chemical bond where a pair of electrons is unequally shared between two atoms.  **10.** a number, figure, symbol, or indicator that is smaller than their normal line of type and is set slightly below or above it.  **11.** the complete transfer of valence electrons between atoms  **13.** a quantity obtained by multiplying quantities together, or from an analogous algebraic operation. |