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CHEMICAL REACTIONS VOCABULARY

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| **Across****2.** are diagrams that show the bonding between atoms of a molecule and the lone pairs of electrons that may exist in the molecule.**5.** a substance that takes part in and undergoes change during a reaction.**12.** a strong force of attraction holding atoms together in a molecule or crystal, resulting from the sharing or transfer of electrons.**14.**  an atom or a molecule in which the total number of electrons is not equal to the total number of protons**15.** when an electron relocates from an atom to another such chemical entity.**16.** are a type of chemical bond where two atoms share a pair of electrons with each other. | **Down****1.** the minimum quantity of energy that the reacting species must possess in order to undergo a specified reaction.**3.** process describes a process or reaction in which the system absorbs energy from its surroundings; usually, but not always, in the form of heat.**4.** a chemical reaction that releases energy by light or heat**6.** a mathematical relationship or rule expressed in symbols.**7.** a chemical bond that involves the sharing of electron pairs between atoms. **8.** the state or power of being reactive or the degree to which a thing is reactive.**9.** a type of chemical bond where a pair of electrons is unequally shared between two atoms.**10.** a number, figure, symbol, or indicator that is smaller than their normal line of type and is set slightly below or above it.**11.** the complete transfer of valence electrons between atoms**13.** a quantity obtained by multiplying quantities together, or from an analogous algebraic operation. |