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Chapter 14

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| **Across**  **2.** gas contribution to the total pressure  **4.** the relationship among pressure, temperature, and volume  **7.** the measure of how much the volume of matter decreases under pressure  **9.** at constant volume and temperature pressure is is equal to partial pressures  **10.** if pressure is constant volume and temperature vary directly  **11.** If volume is constant pressure and temperature are directly proportional  **12.** value of 8.31 (R) | **Down**  **1.** rate of effusion is inversely proportional to the square root of molar mass of the gas  **3.** (P) X (V) = (N) X (R) X (T)  **5.** tendency of molecules moving toward lower concentration  **6.** at constant temperature volume and pressure vary inversely  **8.** gas escaping through a tiny hole |