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Chapter 14

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| **Across****2.** gas contribution to the total pressure**4.** the relationship among pressure, temperature, and volume**7.** the measure of how much the volume of matter decreases under pressure**9.** at constant volume and temperature pressure is is equal to partial pressures**10.** if pressure is constant volume and temperature vary directly **11.** If volume is constant pressure and temperature are directly proportional**12.** value of 8.31 (R) | **Down****1.** rate of effusion is inversely proportional to the square root of molar mass of the gas**3.** (P) X (V) = (N) X (R) X (T) **5.** tendency of molecules moving toward lower concentration**6.** at constant temperature volume and pressure vary inversely **8.** gas escaping through a tiny hole |