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Chapter 6 - Acids, Bases, and Solutions

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| **Across**  **4.** A range of values used to indicate how acidic or basic a substance is; expresses the concentration of hydrogen ions in a solution.  **8.** A substance that tastes bitter, feels slippery, and turns red litmus paper blue.  **12.** An ionic compound made from the neutralization of an acid with a base.  **13.** A mixture containing a solvent and at least one solute that has the same properties throughout; a mixture in which one substance is dissolved in another.  **14.** A substance that tastes sour, reacts with metals and carbonates and turns blue litmus red.  **15.** A mixture containing small, undissolved particles that do not settle out.  **16.** The part of a solution that is usually present in the largest amount and dissolves a solute.  **17.** A measure of how much solute can dissolve in a given solvent at a given temperature.  **18.**  A mixture that has only a little solute dissolved in it. | **Down**  **1.** A mixture in which particles can be seen and easily separated by settling or filtration.  **2.** A mixture that contains as much dissolved solute as is possible at a given temperature.  **3.** A mixture that has a lot of solute dissolved in it.  **5.** A negatively charged ion made of oxygen and hydrogen  **6.** A reaction of an acid with a base, yielding a solution that is not as acidic or basic as the starting solutions were.  **7.** The gradual wearing away of a metal element due to a chemical reaction.  **9.** A positively charged ion formed of a hydrogen atom that has lost it electron.  **10.** The part of a solution that is dissolved by a solvent.  **11.** A compound that changes color in the presence of an acid or a base. |