|  |  |
| --- | --- |
| Name: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ | Date: \_\_\_\_\_\_\_\_\_\_\_\_\_\_ |

Chapter 7 Basics of Chemistry pt. 2

|  |  |
| --- | --- |
| **1.** A substances relative degree of acidity or alkalinity and is measured on a scale of 0 to 14 | **A.** Solutions |
| **2.** An atom or molecule that carries an electrical charge | **B.** Cation |
| **3.** Causes an atom or molecule to split in two, creating a pair of ions with opposite electrical charges | **C.** Redox reactions |
| **4.** An ion with a negative electrical charge | **D.** Reduction |
| **5.** An ion with a positive electrical charge | **E.** Acid mantle |
| **6.** An anion with one oxygen and one hydrogen atom | **F.** Solvent |
| **7.** Substances that have a pH below 7.0, taste sour, turn litmus paper from blue to red | **G.** Anion |
| **8.** Have a pH above 7.0, taste bitter, and turn litmus paper red to blue | **H.** Solute |
| **9.** A protective barrier against certain forms of bacteria and other microorganisms | **I.** Ionization |
| **10.** When an acid is mixed with an alkali in equal proportions to neutralize each other and form water and a salt | **J.** Alkalis |
| **11.** A chemical reaction in which the oxidizing agent is reduced and the reducing agent is oxidized | **K.** Suspensions |
| **12.** The process which oxygen is subtracted from or hydrogen is added to a substance through a chemical reaction | **L.** Acids |
| **13.** Chemical reaction in which the oxidizing agent is reduced and the reducing agent is oxidized | **M.** Immiscible |
| **14.** The rapid oxidation of a substance, accompanied by the production of heat andlight | **N.** Hydroxide |
| **15.** Used to stabilize by preventing oxidation that would otherwise cause a product to turn rancid | **O.** Surfactants |
| **16.** A uniform mixture of two or more mutually miscible substances | **P.** Combustion |
| **17.** Any substance that is dissolved by a solvent to form a solution | **Q.** Emulsion |
| **18.** Any substance that dissolves the solute to form a solution | **R.** Acid-alkali neutralization  |
| **19.** Liquids are mutually soluble | **S.** Antioxidants |
| **20.** Liquids that are not mutually soluble | **T.** Miscible |
| **21.** Unstable mixtures of two or more immiscible substances | **U.** pH (potential hydrogen |
| **22.** Mixtures of two or more immiscible substances united with the aid of an emulsifier | **V.** Oxidation-reduction |
| **23.** Used t emulsify oil and water to create an emulsion | **W.** Ion |
| **24.** One one end of the surfactant molecule that is water loving | **X.** Hydrophilic |
| **25.** One end of the surfactant molecule that is oil loving | **Y.** Lipophilic |