|  |  |
| --- | --- |
| Name: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ | Date: \_\_\_\_\_\_\_\_\_\_\_\_\_\_ |

Chemical Reactions

|  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- |
|  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 1C |  |  |  |  |  |  |  |  |  |  |
|  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  O |  |  |  |  |  |  |  | 2S |  |  |
|  |  |  |  |  | 3E |  |  |  |  |  |  |  |  |  |  |  |  |  |  M |  |  |  |  |  |  |  |  I |  |  |
|  |  |  |  |  |  N |  |  |  |  |  |  |  |  |  |  |  |  |  |  B |  | 4D |  |  |  |  |  |  N |  |  |
|  |  |  |  |  |  D |  |  |  |  |  |  |  |  |  |  |  |  |  |  U |  |  E |  | 5C |  |  |  |  G |  |  |
|  |  |  |  |  |  O |  |  |  | 6R |  |  |  |  |  |  | 7C |  |  |  S |  |  C |  |  O |  |  |  |  L |  |  |
|  |  |  |  |  |  T |  |  |  |  E |  |  |  |  |  |  |  H |  |  |  T |  |  O |  |  E |  |  |  |  E |  |  |
|  |  |  |  |  |  H |  |  |  |  D |  |  |  |  | 8O |  |  E |  |  |  I |  |  M |  |  F |  |  |  |  D |  |  |
|  |  |  |  |  |  E |  |  |  |  U |  |  |  |  |  X |  |  M |  |  |  O |  |  P |  |  F |  |  |  |  I |  |  |
|  |  |  |  | 9P |  R |  O |  D |  U |  C |  T |  |  |  |  I |  |  I |  |  |  N |  |  O |  |  I |  |  |  |  S |  |  |
|  |  |  |  |  |  M |  |  |  |  T |  |  |  |  |  D |  |  C |  |  |  R |  |  S |  |  C |  |  |  |  P |  |  |
|  |  |  |  |  |  I |  |  |  |  I |  |  |  | 10B |  A |  L |  A |  N |  C |  E |  |  I |  |  I |  |  |  |  L |  |  |
|  |  |  |  |  |  C |  |  |  |  O |  | 11C |  |  |  T |  |  L |  |  |  A |  |  T |  |  E |  |  |  |  A |  |  |
|  |  |  |  |  |  R |  | 12S |  Y |  N |  T |  H |  E |  S |  I |  S |  R |  E |  A |  C |  T |  I |  O |  N |  |  |  |  C |  |  |
|  |  |  |  |  |  E |  |  |  |  |  |  E |  |  |  O |  |  E |  |  |  T |  |  O |  |  T |  |  |  |  E |  |  |
| 13R |  E |  A |  C |  T |  A |  N |  T |  |  |  |  M |  |  |  N |  |  A |  |  |  I |  |  N |  |  |  |  |  |  M |  |  |
|  |  |  |  |  |  C |  |  |  |  |  |  I |  |  |  |  |  C |  |  |  O |  |  R |  |  |  |  |  |  E |  |  |
|  |  | 14E |  X |  O |  T |  H |  E |  R |  M |  I |  C |  R |  E |  A |  C |  T |  I |  O |  N |  |  E |  |  |  |  |  |  N |  |  |
|  |  |  |  |  |  I |  |  |  |  |  |  A |  |  |  |  |  I |  |  |  |  |  A |  |  |  |  |  |  T |  |  |
|  |  |  |  |  |  O |  | 15C |  A |  T |  A |  L |  Y |  S |  T |  |  O |  | 16P |  R |  E |  C |  I |  P |  I |  T |  A |  T |  E |  |
|  |  |  |  |  |  N |  |  |  |  |  |  E |  |  |  |  |  N |  |  |  |  |  T |  |  |  |  |  |  E |  |  |
|  |  |  |  |  |  |  |  |  |  |  |  Q |  |  |  |  |  |  |  |  |  |  I |  |  |  |  |  |  A |  |  |
|  |  |  |  |  |  |  |  |  |  | 17S |  U |  B |  S |  C |  R |  I |  P |  T |  |  |  O |  |  |  |  |  |  C |  |  |
|  |  |  |  |  |  |  |  |  |  |  |  A |  |  |  |  |  |  |  |  |  |  N |  |  |  |  |  |  T |  |  |
|  |  |  |  |  |  |  |  |  |  |  |  T |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  I |  |  |
|  |  |  | 18D |  O |  U |  B |  L |  E |  D |  I |  S |  P |  L |  A |  C |  E |  M |  E |  N |  T |  R |  E |  A |  C |  T |  I |  O |  N |  |
|  |  |  |  |  |  |  |  |  |  |  |  O |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  N |  |  |
|  |  |  |  |  |  |  |  |  |  |  |  N |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |

|  |  |
| --- | --- |
| **Across****9.** The new substances that are formed in a chemical reaction.**10.** A condition in which different elements are equal or in the correct proportions. There are three types of visual balance: symmetry, asymmetry, and radial.**12.** A chemical reaction that occurs when a substance reacts with oxygen, releasing energy in the form of heat and light combustion reaction A chemical reaction in which two or more reactants combine to produce a single product.**13.** The starting materials in a chemical reaction.**14.** Energy is released.**15.** A substance that initiates or accelerates a chemical reaction without itself being affected.**16.** A solid that forms from a solution during a chemical reaction.**17.** Tells how many atoms of an element are contained in one molecule or formula unit.**18.** A reaction in which a gas, a solid precipitate, or a molecular compound forms from the exchange of ions between two compounds. | **Down****1.** A chemical reaction that occurs when a substance reacts with oxygen, releasing energy in the form of heat and light.**2.** Chemical reaction in which one element replaces another element in a compound.**3.** Energy is absorbed**4.** A reaction in which a single compound breaks down to form two or more simpler substances.**5.** Tells how many atoms, molecules, or formula units take part in a reaction.**6.** Any process in which electrons are added to an atom or ion.**7.** A process during which chemical bonds between atoms are broken and new ones are formed. Producing one or more different substances.**8.** A chemical change in which a substance combines with oxygen, as when iron oxidizes, forming rust.**11.** A representation of a chemical reaction that uses symbols to show the relationship between the reactants and the products. |