Chemistry Vocab

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| **Across****2.** the point in a titration where the amount of titrant added is enough to completely neutralize the analyte solution**6.** a solution that can resist pH change upon the addition of an acidic or basic components**10.** any solution that has a higher concentration of hydrogen ions than water**14.** an acid that partially dissociates into its ions in an aqueous solution or water**17.** an acid that can donate two proton or hydrogen atom per molecule to an aqueous solution**18.** the moles of an acid or base necessary to change the pH of a solution by 1, divided by the pH change and the volume of buffer in liters**19.** The process by which a water molecule donates a proton to a neighboring water molecule, yielding hydronium and hydroxide ions | **Down****1.** an aqueous solution containing more OH-ions than H+ ions**3.** a scale used to specify how acidic or basic (or alkaline) a water-based solution is**4.** a method of quantitative chemical analysis to determine the concentration of an identified analyte**5.** when an acid and a base react to form water and a salt and involves the combination of H+ ions and OH- ions to generate water**7.** (H3O+)**8.** an acid that is completely dissociated or ionized in an aqueous solution**9.** a mixture of base solids dissolved in water**11.** a chemical base that does not ionize fully in an aqueous solution**12.** the point in a titration at which a reaction is complete**13.** a base that is completely dissociated in an aqueous solution**15.** a solution containing a precisely known concentration of an element or a substance**16.** an acid that has three dissociable protons that undergo stepwise ionization |