Chemistry Vocab

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| **Across**  **2.** the point in a titration where the amount of titrant added is enough to completely neutralize the analyte solution  **6.** a solution that can resist pH change upon the addition of an acidic or basic components  **10.** any solution that has a higher concentration of hydrogen ions than water  **14.** an acid that partially dissociates into its ions in an aqueous solution or water  **17.** an acid that can donate two proton or hydrogen atom per molecule to an aqueous solution  **18.** the moles of an acid or base necessary to change the pH of a solution by 1, divided by the pH change and the volume of buffer in liters  **19.** The process by which a water molecule donates a proton to a neighboring water molecule, yielding hydronium and hydroxide ions | **Down**  **1.** an aqueous solution containing more OH-ions than H+ ions  **3.** a scale used to specify how acidic or basic (or alkaline) a water-based solution is  **4.** a method of quantitative chemical analysis to determine the concentration of an identified analyte  **5.** when an acid and a base react to form water and a salt and involves the combination of H+ ions and OH- ions to generate water  **7.** (H3O+)  **8.** an acid that is completely dissociated or ionized in an aqueous solution  **9.** a mixture of base solids dissolved in water  **11.** a chemical base that does not ionize fully in an aqueous solution  **12.** the point in a titration at which a reaction is complete  **13.** a base that is completely dissociated in an aqueous solution  **15.** a solution containing a precisely known concentration of an element or a substance  **16.** an acid that has three dissociable protons that undergo stepwise ionization |