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| Name: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ | Date: \_\_\_\_\_\_\_\_\_ | Period: \_\_\_\_\_\_\_ |

Chemistry Vocab

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|  |  |  |  |  |  |  |  |  |  |  | 1  N |  |  |  |  |  |  |  |  |  | 2  P |  |  |  |  |  |  |  |  |
|  |  |  |  |  |  |  |  |  |  | 3  R | E | A | C | T | 4  I | V | I | T | Y |  | E |  |  |  |  |  |  |  |  |
|  |  |  |  |  |  |  |  |  |  |  | G |  |  |  | O |  |  |  |  |  | R |  |  |  |  |  |  |  |  |
|  |  |  |  | 5  A |  |  |  |  | 6  N |  | A |  |  |  | N |  |  |  |  |  | I |  |  |  |  |  |  |  |  |
|  |  |  |  | L |  |  |  |  | O |  | T |  |  |  |  |  | 7  P | R | O | T | O | N |  |  |  |  |  |  |  |
|  |  |  |  | K |  |  |  |  | B |  | I |  | 8  M |  | 9  P |  |  |  |  |  | D |  |  |  |  |  |  |  |  |
|  |  |  |  | A |  |  |  |  | L |  | V |  | A |  | H |  |  |  |  |  | I |  |  |  |  |  |  |  |  |
|  |  |  |  | L |  |  | 10  E | L | E | M | E | N | T |  | Y |  |  |  |  |  | C |  |  |  |  |  | 11  P |  |  |
|  |  |  |  | I |  |  |  |  | G |  |  |  | T |  | S |  |  |  |  |  | T |  |  |  |  |  | O |  |  |
|  |  |  |  | N |  |  |  |  | A |  |  | 12  P | E | R | I | O | D | S |  |  | A |  |  | 13  M |  |  | S |  |  |
|  |  |  |  | E |  |  |  |  | S |  |  |  | R |  | C |  |  |  |  |  | B |  |  | A |  |  | I |  |  |
|  |  |  |  | E |  |  |  |  |  |  |  |  |  |  | A |  | 14  A |  |  |  | L |  |  | L |  |  | T |  |  |
|  |  |  |  | A |  |  |  |  |  | 15  D | U | C | T | I | L | I | T | Y |  | 16  M | E | T | A | L | L | O | I | D | S |
|  |  |  |  | R |  |  |  |  |  |  |  |  |  |  | C |  | O |  |  |  |  |  |  | E |  |  | V |  |  |
|  |  |  |  | T |  |  |  |  |  |  |  |  |  | 17  C | H | E | M | I | C | A | L | C | H | A | N | G | E | S |  |
|  |  |  |  | H |  |  |  |  | 18  F |  |  |  |  |  | A |  |  |  |  |  |  |  |  | B |  |  |  |  |  |
|  |  |  |  | M |  |  |  | 19  E | L | E | C | T | R | O | N |  |  |  |  |  |  |  |  | I |  |  |  |  |  |
|  |  |  |  | E |  |  |  |  | A |  |  |  |  |  | G |  |  |  |  |  |  |  |  | L |  |  |  |  |  |
|  |  |  |  | T |  |  | 20  F | A | M | I | L | Y |  | 21  N | E | U | T | R | O | N |  |  |  | I |  |  |  |  |  |
|  |  |  |  | A |  |  |  |  | M |  |  |  |  |  | S |  |  |  |  |  |  |  |  | T |  |  |  |  |  |
| 22  A | L | K | A | L | I | M | E | T | A | L | S |  |  |  |  |  |  |  |  |  |  |  |  | Y |  |  |  |  |  |
|  |  |  |  | S |  |  |  |  | B |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |
|  |  |  |  |  | 23  C | H | E | M | I | C | A | L | P | R | O | P | E | R | T | I | E | S |  |  |  |  |  |  |  |
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|  |  |  |  |  | 24  P | H | Y | S | I | C | A | L | P | R | O | P | E | R | T | I | E | S |  |  |  |  |  |  |  |
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| **Across**  **3.** shows the materials's ability to react with another substance  **7.**  the positive subatomic particle in the nucleus which determines its atomic number and along with the neutron determines the atomic mass  **10.**  one of a class of substances that cannot be separated into simpler substances by chemical means.  **12.**  the columns of the periodic table; they are grouped by similar chemical and physical properties  **15.**  shows the material's ability to be drawn into a thin wire  **16.**  a group of elements consisting of properties of metals and nonmetals and are located on the diagonal line  **17.**  A usually irreversible change that changes the arrangement of the atoms in the chemical composition that involves a formation of a new substance.  **19.**  the particle in the surrounding cloud of the nucleus having a negative charge and is 1000 times smaller than a proton  **20.**  the rows of the periodic table; they go in an ascending order with the atomic number  **21.**  the particle in the nucleus of the atom which has the biggest mass and no charge which determines the atomic mass with the proton  **22.**  the most reactive metal group located at the leftmost of the the periodic table.  **23.**  a property or characteristic of a substance that is observed during a reaction in which the chemical composition or identity of the substance is changed  **24.**  a property of a substance of matter that can be observed without changing the chemical composition of the substance. | **Down**  **1.**  the charge of an electron  **2.**  a set of elements which are organized on a table according to the atomic number, electron configuration, and physical/chemical properties  **4.**  an atom that is either mostly negatively charged or mostly positively charged  **5.**  any of the group of bivalent metals including barium, radium, strontium, calcium, and, usually, magnesium, the hydroxides of which are alkali's but less soluble than those of the alkali metals.  **6.**  the mostly nonreactive group which are at the rightmost of the periodic table  **8.**  anything that has mass and takes up space  **9.**  A usually reversible change in the physical properties of a substance.  **11.**  the charge of a proton  **13.**  show the material's ability to be molded into a different shape  **14.** What is the smallest particle of an element consisting the same chemical properties of the element  **18.**  shows the materials's ability to catch fire |