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Chemistry crossword

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| **Across****2.** \_\_\_\_ \_\_\_\_ \_\_\_\_ states: mass/matter is neither created nor destroyed. This explains why the mass of a substance remains constant during phase changes**8.** Two or more atoms joined together.**11.** All matter in the universe is composed of approximately \_\_\_ atoms**13.** \_\_\_\_ \_\_\_\_ is the measurement of the motion of the atoms and molecules. It is the energy the substance has related to its temperature**16.** \_\_\_\_ \_\_\_\_ can be used to identify substances**17.** \_\_\_\_\_\_\_\_ are a group of substances composed of a single kind of atom**18.** The amount of matter in an object**19.** \_\_\_\_ and its properties are determined by the arrangement of atoms**20.** Materials or \_\_\_\_ composed of 2 or more substances are in the same place, but do not chemically bond | **Down****1.** Shape, conductivity, and texture are examples of \_\_\_\_\_\_\_\_\_\_\_\_\_**3.** \_\_\_\_ \_\_\_\_ have a repeated pattern and a distinctive melting point**4.** The three states of matter**5.** The \_\_\_\_ of iron shavings and styrofoam will separate with magnetic attraction**6.** When a bow is pulled back or a spring is pushed down, either will gain \_\_\_\_ \_\_\_\_**7.** Energy stored in bonds of atoms and molecules**9.** The amount of space matter occupies**10.** This determines a substances state of matter (solid, liquid or gas)**12.** The compactness of a material. A material's \_\_\_\_\_\_ will not change depending on how much of the material you have.**14.** \_\_\_\_\_ have mass, take up space and are in constant motion**15.** Consisting of only one kind of atom, can be broken down into the same substance, and exists as either atoms or molecules are properties of \_\_\_\_\_\_\_\_ |

   Atoms       Molecules       Elements       100       Elements       Density       Mixture       Physical properties of matter       Volume       Mass       Temperature       Matter       Mixtures       Chemical and physical properties       Thermal energy       Crystalline solids       solid liquid gas       Law of conservation       Chemical Energy       Potential Energy