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Ionic Compounds

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| **Across**  **1.** a force that holds two atoms together  **3.** An aqueous solution of an ionic compound that conducts electricity  **5.** The electrons involved in metallic bonding that are free to move easily from one atom to the next throughout the metal and are not attached to a particular atom  **13.** positively charged ion  **14.** Compound with an ionic bond  **15.** The energy required to separate 1 mole of the ions of an ionic compound  **16.** An ion made up of two or more atoms bonded together that acts as a single unit with a net charge | **Down**  **2.** Proposes that all metal atoms in a metallic solid contribute their valence electrons to form a "sea" of electrons  **4.** The attraction of a metallic cation for delocalized electrons  **6.** A mixture of elements that has metallic properties  **7.** an ion formed by only 1 atom  **8.** a force that holds oppositely charged particles together  **9.** The simplest ratio of ions represented in an ionic compound  **10.** A 3-D arrangement of particles so that all positives are surrounded by negatives and all negatives are surrounded by positives  **11.** A polyatomic ion composed of an element, usually a nonmetal, bonded to one or more oxygen atoms  **12.** negatively charged ion |