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Ionic Compounds

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| **Across****1.** a force that holds two atoms together**3.** An aqueous solution of an ionic compound that conducts electricity**5.** The electrons involved in metallic bonding that are free to move easily from one atom to the next throughout the metal and are not attached to a particular atom**13.** positively charged ion**14.** Compound with an ionic bond**15.** The energy required to separate 1 mole of the ions of an ionic compound**16.** An ion made up of two or more atoms bonded together that acts as a single unit with a net charge | **Down****2.** Proposes that all metal atoms in a metallic solid contribute their valence electrons to form a "sea" of electrons**4.** The attraction of a metallic cation for delocalized electrons**6.** A mixture of elements that has metallic properties**7.** an ion formed by only 1 atom**8.** a force that holds oppositely charged particles together**9.** The simplest ratio of ions represented in an ionic compound**10.** A 3-D arrangement of particles so that all positives are surrounded by negatives and all negatives are surrounded by positives**11.** A polyatomic ion composed of an element, usually a nonmetal, bonded to one or more oxygen atoms**12.** negatively charged ion |