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Molarity and Dilutions

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| **Across****5.** the erratic movement of colloid particles**7.** the overall energy change that occurs during the solution formation process**10.** two liquids that can be mixed together but separate shortly after**12.** heterogeneous mixture of intermediate sized particles**13.** two liquids that are soluble in each other in any proportion**15.** contains the maximum amount of dissolved solute for a given amount of solvent at a specific temperature and pressure**16.** a substance that does not dissolve in a solvent**17.** when dispersed colloid particles scatter light | **Down****1.** a mixture containing particles that settle out if left undisturbed**2.** the ratio of the number of moles of solute or solvent in solution to the total number of moles of solute and solvent**3.** contains less dissolved solute for a given temperate and pressure**4.** contained more dissolved solute than a saturated solution at the same temperature**6.** the ratio of the number of moles of solute dissolved in 1 kg of solvent**8.** measure of how much solute is dissolved in a specific amount of solvent or solution**9.** the number of moles of solute dissolved per liter of solution**11.** is able to be dissolved into a solvent**14.** the process of surrounding solute particles with solvent particles to form a solution |