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Molarity and Dilutions

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| **Across**  **5.** the erratic movement of colloid particles  **7.** the overall energy change that occurs during the solution formation process  **10.** two liquids that can be mixed together but separate shortly after  **12.** heterogeneous mixture of intermediate sized particles  **13.** two liquids that are soluble in each other in any proportion  **15.** contains the maximum amount of dissolved solute for a given amount of solvent at a specific temperature and pressure  **16.** a substance that does not dissolve in a solvent  **17.** when dispersed colloid particles scatter light | **Down**  **1.** a mixture containing particles that settle out if left undisturbed  **2.** the ratio of the number of moles of solute or solvent in solution to the total number of moles of solute and solvent  **3.** contains less dissolved solute for a given temperate and pressure  **4.** contained more dissolved solute than a saturated solution at the same temperature  **6.** the ratio of the number of moles of solute dissolved in 1 kg of solvent  **8.** measure of how much solute is dissolved in a specific amount of solvent or solution  **9.** the number of moles of solute dissolved per liter of solution  **11.** is able to be dissolved into a solvent  **14.** the process of surrounding solute particles with solvent particles to form a solution |