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| Name: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ |

Mole Day: The Periodic Table of the EleMOLEments

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|  |  |  |  |  |  | 1  A |  |  |  | 2  I |  |  |  |  |  |  |  |  |  |
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|  |  |  |  |  |  | A |  |  | 3  C | A | R | B | O | N | 4  1 | 2 |  |  |  |
|  |  |  |  |  |  | D |  | 5  P |  | L |  |  |  |  | 8 |  |  |  |  |
|  |  |  |  |  |  | E |  | R |  | Y |  |  |  |  | 6 |  |  |  |  |
|  |  |  |  |  | 6  M | O | L | E |  |  |  | 7  6 |  |  | 0 |  | 8  C |  |  |
|  |  |  |  |  |  |  |  | T |  | 9  M |  | 0 |  |  |  |  | A |  |  |
|  |  | 10  1 |  |  |  | 11  G |  | T |  | A |  | 2 |  |  | 12  1 |  | N |  |  |
|  |  | 8 |  |  |  | R |  | Y |  | S |  | 2 |  |  | 2 |  | I |  |  |
|  |  | 1 |  |  |  | A |  | D |  | S |  | X |  |  | 0 |  | Z |  |  |
|  |  | 1 |  | 13  C | O | M | P | A | R | E |  | 14  1 | 2 | 0 | 1 |  | Z |  |  |
|  |  |  |  |  |  | S |  | R |  | S |  | 0 |  |  |  |  | A |  |  |
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|  |  |  |  | 15  O | C | T | O | B | E | R | 2 | 3 |  |  |  | 16  M | O | L | E |
|  |  |  |  |  |  |  |  | I |  |  |  |  |  |  |  |  |  |  |  |
|  | 17  A | N | Y | T | H | I | N | G |  |  |  |  |  |  |  |  |  |  |  |
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| **Across**  **3.** Atom decided upon in First Chemical Congress  **6.** A small brown animal that burrows underground  **13.** Allows chemists to \_\_\_\_\_\_\_\_\_ the weight of substances given they have the same number of atoms  **14.** Mass of 1 atom of carbon in amu  **15.** National Mole Day  **16.** The amount of substance containing the same number of atoms as 12 grams of carbon-12  **17.** You can have a mole of... | **Down**  **1.** Avogadro's first name  **2.** Country Avogadro is from  **4.** Year of the First Chemical Congress  **5.** How big the number 6.022 x 10^23 is  **7.** Avogadro's Number  **8.** Presented paper in support of Avogadro's hypothesis  **9.** Purpose of First Chemical Congress was to determine how \_\_\_\_\_\_\_\_ of different elements could be measured  **10.** When Avogadro first proposed his hypothesis  **11.** Helps convert from \_\_\_\_\_\_\_ to atoms  **12.** Mass of 6.022 x 10^23 atoms in grams for carbon |