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| Name: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ | Date: \_\_\_\_\_\_\_\_\_ | Period: \_\_\_\_\_\_\_ |

Oxidation-Reduction Vocab

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| **Across**  **5.** Electrochemical cell in which a spontaneous chemical reaction causes a flow of electrons.  **6.** Loss of electrons and an increase in \_\_\_\_\_\_\_\_\_ state.  **7.** Process in which an electric current forces a nonspontaneous reaction to occur.  **8.** Gain of electrons and the loss of oxidation number.  **9.** System in which there's an electric current flowing while a chemical reaction occurs.  **10.** Site at which oxidation or reduction; an anode or cathode.  **11.** Reaction that shows either the oxidation or reduction portion of a redox reaction. | **Down**  **1.** An oxidation-reduction reaction.  **2.** Cell that requires a nonspontaneous chemical reaction to occur (uses battery).  **3.** Number assigned to keep track of electron gain or loss in redox reactions.  **4.** Part of a voltaic cell that connects two containers and allows flow of ions. |

   electrolytic cell       half-reaction       salt bridge       electrochemical cell       electrode       oxidation       electrolysis       redox       reduction       oxidation number       voltaic cell