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| Name: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ |

Periodic Table And Atoms

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|  |  |  |  |  |  |  |  | 2  V | A | L | A | N | C | E | E | L | E | C | T | R | O | N | S |  |  |  |  |  |  |
|  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | V |  |  |  |  |  |  |  |  |
|  |  |  |  |  |  |  |  | 3  E | N | D | O | T | H | E | R | M | I | C | R | E | A | C | T | I | O | N |  |  |  |
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|  |  |  |  |  |  |  |  |  | 5  I |  |  | 6  E |  |  |  |  |  |  |  |  | T |  |  |  |  |  |  |  |  |
|  |  |  |  |  |  |  |  |  | O |  |  | L |  |  |  |  |  |  | 7  P |  | B |  |  |  |  |  |  |  |  |
|  |  |  |  | 8  P | R | O | T | O | N | S |  | E |  |  |  |  |  |  | E |  | O |  |  | 9  R |  |  |  | 10  A |  |
|  |  |  |  |  |  |  |  |  | I |  |  | C |  |  |  |  | 11  M |  | R |  | N |  |  | A |  |  |  | V |  |
|  |  |  |  |  | 12  I |  |  |  | C |  |  | T |  | 13  N |  | 14  P | E | R | I | O | D | I | C | T | A | B | L | E |  |
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|  |  |  |  |  | S |  |  | 15  C | O | M | P | O | U | N | D |  | A |  | D |  |  |  | 16  I | O | N |  |  | A |  |
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|  |  |  |  |  | A |  |  |  |  |  |  |  |  | T |  |  | O |  |  |  |  |  |  | E |  |  |  | A |  |
|  |  | 18  E | X | O | T | H | E | R | M | I | C | R | E | A | C | T | I | O | N |  | 19  C | A | T | A | L | Y | S | T |  |
|  |  |  |  |  | O |  |  |  |  |  |  |  |  | L |  |  | D |  |  |  |  |  |  | C |  |  |  | O |  |
|  |  |  |  |  | R |  | 20  E | L | E | M | E | 21  N | T | S |  |  | S |  |  |  |  |  |  | T |  |  |  | M |  |
|  |  |  |  |  |  |  |  |  |  |  |  | E |  |  |  |  |  |  | 22  G |  |  |  |  | I |  |  |  | I |  |
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|  | 24  A | T | O | M | I | C | N | U | M | B | E | R |  | 25  M | A | S | S | N | U | M | B | E | R |  |  |  |  | A |  |
|  |  |  |  |  |  |  |  |  |  |  |  | O |  |  |  |  |  |  | P |  |  |  |  |  |  |  |  | S |  |
|  |  |  |  |  |  |  |  |  |  | 26  C | O | N | D | U | C | T | O | R | S |  | 27  I | S | O | T | O | P | E | S |  |

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| **Across**  **2.** The electrons in the outermost energy level of Niels Bohr's model of an atom  **3.** Chemical Reactions that absorb energy  **4.** The positively charged central part of an atom  **8.** A positively charged particles that make up a nucleus  **14.** A chart of elements arranged into rows and columns based on their chemical and physical properties  **15.** A substance that is made of two or more elements chemically joined in a specific combination  **16.** An atom no longer neutrally charged because it has lost electrons  **17.** A group located on the left side of the Periodic Table that is a good conductor and shine when polished  **18.** A chemical reaction that releases heat of light energy  **19.** A substance that speeds up chemical reactions  **20.** Pure substances that cannot be broken down into any simpler substances  **23.** An area around an atomic nucleus where an electron is most likely to be found  **24.** The number of protons in an atom of an element  **25.** The sum of the number of protons and neutrons in an atom  **26.** An object that has low electrical resistance and can allow electricity to flow easily  **27.** Atoms of the same element that have different numbers of neutrons | **Down**  **1.** A bond formed when two atoms share valence electrons  **5.** An attraction that holds ions close together  **6.** A particle with a single negative charge  **7.** Horizontal rows on a Periodic Table  **9.** The speed at which a reaction occurs  **10.** The average mass of an element's isotopes  **11.** Elements in the middle of the Periodic Table and are semiconductors.They also have properties that are similar to both metals and nonmetals  **12.** An object that has higher electrical resistance and prevents electricity easily through a material  **13.** Another group located on the right side of the Periodic Table that are insulators and are not shiny  **21.** A nuetral particle in the nucleus of an atom  **22.** Vertical columns on the periodic table |