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Periodic Table Crossword Puzzle

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| **Across**  **4.** any of the elements lithium, sodium, potassium, rubidium, cesium, and francium, occupying Group IA (1) of the periodic table. They are very reactive, electropositive, monovalent metals forming strongly alkaline hydroxides.  **5.** is of an atom or molecule is defined as the amount of energy released or spent when an electron is added to a neutral atom or molecule in the gaseous state to form a negative ion.  **7.** arising from the arrangement of the periodic table, provide chemists with an invaluable tool to quickly predict an element's properties.  **8.** is the process by which an atom or a molecule acquires a negative or positive charge by gaining or losing electrons to form ions, often in conjunction with other chemical changes.  **9.** a chemical element (as boron, carbon, or nitrogen) that lacks the characteristics of a metal and that is able to form anions, acidic oxides, acids, and stable compounds with hydrogen.  **10.**  is an element that doesn't have the characteristics of metal including: ability to conduct heat or electricity, luster, or flexibility. An example of a nonmetal element is carbon. | **Down**  **1.** introduced by Niels Bohr and Ernest Rutherford in 1913, depicts the atom as a small, positively charged nucleus surrounded by electrons  **2.** any of the six highly reactive chemical elements in group IIA of the periodic table including, in order of increasing reactivity, beryllium, magnesium, calcium, strontium, barium, and radium: an oxide of any of these metals is called  **3.** is a measure of the tendency of an atom to attract a bonding pair of electrons.  **6.** any of the gaseous elements helium, neon, argon, krypton, xenon, and radon, occupying Group 0 (18) of the periodic table. They were long believed to be totally unreactive but compounds of xenon, krypton, and radon are now known. |