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Periodic Table Vocabulary Crossword

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|  |  |  |  |  |  | 9N |  O |  N |  M |  E |  T |  A |  L |  |  L |  | 10P |  E |  R |  I |  O |  D |  |  I |  |  |  |  |  |
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| **Across****5.** group 18 elements on the periodic table that contain 8 valence electrons (He has 2) and a full valence shell making them very stable and inert.**8.** describes how likely an element is to form bonds with other elements.**9.** an element that does not conduct electricity or heat and is usually a gas at room temperature. Nonmetals are brittle, have high ionization energies and high electronegativity values. Nonmetals tend to gain electrons to form anions**10.** a horizontal row (left to right) in the periodic table.**13.** a chart that organizes information about all of the known elements according to their atomic number.**14.**  group 1 metals on the periodic table that contain 1 valence electron and lose their valence electrons the most easily, making them the most reactive metals.**16.**  the smallest unit of an element that has all of the properties of the element; basic building block of matter.**17.** the number of protons and neutrons in the nucleus of one atom of the element.**19.** the distance between the nucleus of an atom and it's outermost energy level**20.** group 17 nonmetals on the periodic table that contain 7 valence electrons. They only need to gain 1 valence electron to have a stable octet. They gain valence electrons the most readily, making them the most reactive nonmetals. | **Down****1.**  group 1 metals on the periodic table that contain 2 valence electrons and are the second most reactive metals.**2.** the attraction a nucleus has resulting from its number of protons **3.** an element that has some properties of a metal and some properties of a nonmetal. The metalloids are found on the boron staircase, there are 7 metalloids: B, Si, Ge, As, Sb, Te, and Po.**4.**  group 3-­12 on the periodic table. They have varying valence electrons and do not follow the normal trends of the other metals. They form brightly colored compounds and ions in solution.**6.**  a vertical column (up and down) on the periodic table.**7.** a substance that cannot be broken down into a simpler substance by ordinary chemical means**11.** the charge of an atoms nucleus resulting from its number of protons**12.**  the number of protons contained in each nucleus of its atoms of the element.**15.** elements and/or compounds that when put together are unable to react chemically. The noble gases (group 18) elements are inert because of a full valence shell.**18.** an element or substance that conducts heat and electricity, is malleable and ductile and has low ionization energy and low electronegativity values. Metals |