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Regents Chemistry Chapter 12 Solution Vocabulary

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| 13P |  E |  R |  C |  E |  N |  T |  C |  O |  M |  P |  O |  S |  I |  T |  I |  O |  N |  |  O |  | 14A |  L |  L |  O |  Y |  |  |  |  |
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|  |  | 22S |  U |  P |  E |  R |  S |  A |  T |  U |  R |  A |  T |  E |  D |  |  E |  |  |  | 23I |  N |  S |  O |  L |  U |  B |  L |  E |
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| **Across****3.** the boiling point of a solution is higher than the expected boiling point of the pure solvent (colligative property)**6.** A heterogeneous mixture in which relatively large particles are suspended in a liquid**7.** having a relatively low concentration of solute in a mixture**13.** % comp = (part/whole) x 100**14.** a homogenous mixture/solution containing at least one metal. Ex: brass, steel, bronze**16.**  A sample of matter consisting of more than one pure substance and more than one phase**22.** a solution in which the concentration of solute ishigher than the solubility; more solute is dissolved than should be under a given set of conditions**23.** Refers to a substance that does not dissolve in a solvent to any siginifcant degree**24.** the temperature at which a liquid undergoes a phase change form liquid to gas; the temperature at which the vapor pressure of a liquid is equal to the atmospheric pressure.**25.** A substance dissolved in a solvent to make a solution **26.** A sample of matter consisting of more than one pure substance with properties that do not vary within the sample**27.** A solution with a concentration lower than it's equilibrium solubility; a solution in which more solute can be dissolved  | **Down****1.** An insoluble substance that has formed from a chemical reaction between substances dissolved in a soltuion **2.** two or more pure substances PHYSICALLY combined; a combination of two or more pure substances that can be sperated by physical means**4.** the freezing point/melting point of a solution is lower than the freezing point/melting point of the pure solvent (colligative property)**5.** The most abundant component in a solution **8.** Having a relatively large amount of substance present in a unit amount of mixture. For example, a 12 M HCL solution is more concentrated than an 0.001 M HCL solution**9.** a homogeneous mixture**10.** a heterogeneous mixture composed of tiny particles suspended in another material. The particles are larger than the particles in a solution but smaller than particles in a suspension. Ex; milk, blood**11.** a measure of concentration; ppm = parts of solute/million parts of solution **12.** a measure of concentration; M=moles of solute/liters of solution**15.** capable of being dissolved ina solvent**17.** A measure of the amount of solute present in a unit amont of mixture. (Ex: ppm = parts per million, molarity = moles solute/liter solution); the process of increasing the amount of substance in a given amount of mixture**18.** a meausre of the concentration of a substance in a saturated solution; a measure of how much a substance can dissolve in a given amount of solvent**19.** Two liquids are considered "miscible" or mixable if shaking them together results in a single liquid phase with no visible seperation**20.** a solution that has reached equilibrium; a solution which can't dissolve any more solute**21.** a homogeneous mixture/solution in whixh a solute is dissloved in water |

   Alloy       Aqueous       Boiling Point       Dilute       Molarity       Insoluble       Miscible       Mixture       Colloid       Boiling Point Elevation        Concentrated        Concentration       Freezing Point Depression       Parts Per Million       Homogeneous        Heterogeneous       Solubility       Unsaturated       Suspension       Supersaturated       Percent Composition        Precipitate       Saturated       Soluble       Solution       Solvent       Solute