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Science Crossword For Curriculum Day

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| **Across****4.**  A substance that increases the rate of a chemical reaction. It is not used or changed in the reaction.**8.**  A saturated hydrocarbon that has no double bonds between the carbon atoms**11.**  A tiny particle with a negative charge. Electrons orbit the nucleus in atoms or ions.**12.**  An explanation of chemical reactions that says that particles colliding have to have a sufficient amount of energy in order for a reaction to take place.**14.** A huge 3D network of atoms or ions (e.g. the giant ionic lattice in sodium chloride).**19.** The attraction between two atoms that share one or more pairs of electrons.**20.**  A number that shows how strongly acidic or alkaline a solution is.**21.** A tiny particle with a negative charge. Electrons orbit the nucleus in atoms or ions.**22.** A substance we start with before a chemical reaction takes place.**23.** The electrostatic force of attraction between positively and negatively charged ions**24.**  The amount of product obtained in a reaction.**25.** The average mass of the atom of an element compared with carbon-12 (which is given a mass of exactly 12). The average mass must take into account the proportions of the naturally occurring isotopes of the element. | **Down****1.** A dense particle found in the nucleus of an atom. It is electrically neutral. They have no overall charge.**2.** The total number of protons and neutrons in an atom.**3.** Easily ignited and capable of burning rapidly.**5.**  Bonding electron that is no longer associated with any one particular atom.**6.** A huge 3D network of covalently bonded atoms (e.g. the giant lattice of carbon atoms in diamond or graphite).**7.**  A mixture of metals (and sometimes non metals)**9.**  Atom that has the same number of protons but a different number of neutrons. It has the same atomic number but a different mass number**10.** The reaction when oxygen is added to a substance (or when electrons are lost).**13.** A reaction that takes in energy from the surroundings**15.**  The minimum energy needed to start a reaction**16.**  The amount of substance in the relative atomic or formula mass of a substance in grams.**17.**  A method for measuring the volumes of two solutions that react together.**18.**  A reaction that gives out energy to the surroundings. |