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Science worksheet

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| **Across**  **3.** The combination of three gas laws (Boyles,Charles and Gay Lussac)  **7.** Application of Boyles Law  **8.** amount of space/gas occupied  **9.** Unit of pressure  **10.** a scientific instrument used to measure air pressure.  **12.** measure of KE of the particles  **13.** states that P X V = n X (R) X T, where P is pressure, V is volume, n is the number of moles of molecules, T is the absolute temperature, and R is the gas constant (8.314 joules per degree Kelvin or 1.985 calories per degree Celsius).  **15.** Law that describe the relationship between temperature and pressure | **Down**  **1.** effect of forces of a colliding molecules  **2.** states the volume is inversely proportional to pressure at constant temperature  **4.** Unit of volume  **5.** law that tends to describe how volume change when temperature increase or decrease  **6.** a standard scientific unit for measuring large quantities of very small entities such as atoms, molecules, or other specified particles.  **11.** is the primary unit of temperature measurement in the physical sciences, but is often used in conjunction with the degree Celsius  **14.** states that the volume of gas is directly proportional to the number of moles |