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| Name: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ | Date: \_\_\_\_\_\_\_\_\_ | Period: \_\_\_\_\_\_\_ |

Solubility crossword

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|  |  |  |  |  |  L |  |  |  |  |  |  |  |  |  |  |  | 3L |  |  |
|  |  |  |  |  |  A |  |  |  |  |  |  |  |  |  |  |  |  I |  |  |
|  |  |  |  |  |  R |  | 4S |  |  |  |  |  |  |  |  |  |  K |  |  |
|  |  | 5D |  | 6D |  I |  L |  U |  T |  E |  | 7S |  O |  L |  U |  B |  L |  E |  |  |
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| **Across****2.** A homogenous mixture consisting of a solute dissolved in a solvent; can be any combination of solids, liquids, and/or gases**6.** Describes a solution that has a relatively small amount of dissolved solute**7.** Can be dissolved by a particular solvent**11.** Cannot be dissolved by a particular solvent**12.** A graph showing the solubility of a substance at various temperatures**13.** In a solution, this is the substance being dissolved**14.** A solution that contains less solute than it is capable of dissolving at a given temperature | **Down****1.** Exists when a molecule has a clustering of negative charge on one side due to unequal sharing of electrons**3.** Nonpolar solvents can only dissolve nonpolar solutes; Polar solvents can only dissolve polar or ionic solutes**4.** A solution that contains more solute than it is capable of dissolving at a given temperature**5.** The process of adding more solvent to a solution in order to make it less concentrated**8.** The # of moles of solute in 1.0 liter of solution**9.** Solvent particles surround solute paricles**10.** In a solution, this is the substance doing the dissolving |

   Solute       Solution       Dilute       Solubilitycurve       Insoluble       Soluble       Solvation       Insoluble       Like Dissolves Like       Unsaturated       Polarity       Dilution        Molarity       Supersaturated