States of matter

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| **Across****3.** All matter naturally exists on earth most commonly as a solid, a liquid, or a gas**4.** Doesn’t have a uniform composition. Individual substance remain distinct **11.** Used a porous barrier to separate a solid from a liquid **14.** Chemical combination of 2 or more different elements **15.** Remains the same no matter how much of a substance is present **16.** Form of matter that flows to conform to the shape of its containers entire volume, and its easily compressed **17.** Own definite shape and volume is in compressible, and expands only slightly when heated **18.** Form of matter that flows, has constant volume, takes the shape of its container **19.** Involving one or more substances changing into new substance also called chemical reaction **20.** Alters the physical properties of a substance but does not change its composition  | **Down****1.** Gaseous state of a substance that is a liquid or a solid at room temperature **2.** States that regardless of the amount a compound of the same elements in the same portion by mass**5.** Has a uniform and unchanging composition**6.** Separation technique that produces pure solid particles of a substance from a solution that contains the dissolved substance **7.** Another word for solution **8.** Physical properties such as mass, length, and volume **9.** Physically separate most homogeneous mixture based on the differences in the boiling points of the substances involved **10.** Used to separate the components of a mixture based on the tendency of each component to travel or be drawn across the surface of another material **12.** Substance that can not be broken down in simpler substances by physical or chemical means**13.** Uniform mixture that may contain solids, liquids, or gas |

   Homogeneous mixture        Compound        Intensive properties        Physical change        Elements       Gas       Crystallization        Distillation        Liquid        Vapor       Law of definite proportions        Substance        Solid        Solution        Chemical change        Chromatography        Extensive properties        Heterogeneous mixture        Filtration        State of matter