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| Name: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ | Date: \_\_\_\_\_\_\_\_\_ | Period: \_\_\_\_\_\_\_ |

Structure of the Atom

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|  |  |  |  |  |  |  |  |  |  | 2C |  |  |  |  |  |  |  |  |  | 3S |  |  M |  |  |  |  |  |  |  |
|  |  |  |  |  |  |  |  |  |  |  O |  |  |  |  |  |  | 4V |  |  |  Y |  |  E |  |  |  |  |  |  |  |
|  |  |  |  |  |  |  |  |  |  |  M |  | 5A |  |  |  |  |  A |  |  | 6M |  A |  N |  |  |  |  |  |  |  |
|  |  |  |  |  |  |  |  |  |  |  P |  |  T |  |  |  |  |  L |  |  |  B |  |  T |  |  |  |  |  |  |  |
|  |  |  |  |  |  |  |  |  |  |  O |  |  O |  |  |  | 7P |  E |  R |  I |  O |  D |  S |  | 8E |  |  |  |  |  |
|  |  |  |  |  |  |  |  |  |  |  U |  |  M |  |  |  |  |  N |  |  |  L |  |  |  |  L |  |  |  |  |  |
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|  |  |  |  |  | 9P |  E |  R |  I |  O |  D |  I |  C |  T |  A |  B |  L |  E |  |  | 10N |  | 11E |  |  C |  |  |  |  |  |
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|  |  |  |  |  |  |  | 12A |  |  | 13A |  |  A |  |  | 14M |  O |  L |  E |  C |  U |  L |  E |  |  R |  |  |  |  |  |
|  |  |  |  |  |  |  |  M |  |  |  P |  |  S |  | 15A |  |  |  E |  |  |  T |  |  C |  |  O |  |  |  |  |  |
|  |  |  |  |  |  | 16N |  U |  C |  L |  E |  U |  S |  |  T |  |  |  C |  | 17P |  R |  O |  T |  O |  N |  S |  |  |  |  |
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|  |  |  |  |  |  |  |  | 18M |  A |  S |  S |  N |  U |  M |  B |  E |  R |  |  |  N |  |  O |  |  L |  |  |  |  |  |
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| **Across****6.** Mass Number - Atomic Number = Number of Neutrons**7.** The horizontal rows in the periodic table**9.** Arrangement of elements according to atomic number on a table**14.** A single atom or several atoms bound together electro magnetically bonded.**16.** Center of atom**17.** Positive particles around the nucleus**18.** The sum of the number of neutrons and protons**20.**  1. All matter consists of minuscule particles called atoms. 2. All atoms of a given element are identical to each other. 3. All atoms of a given element are different than those of other elements. 4. Atoms of one element combine with other elements to create compounds. They always combine in equal amounts. | **Down****1.** Pure substances that can't be broken down or changed**2.** 2 or more elements chemically bonded**3.** Abbreviation for each element**4.** The electrons on the very outer ring of the atom**5.** The property of an atom that causes it to have weight**8.** Located outside the nucleus with electrons**10.** Neutral charged particles around the nucleus**11.** Negative particles outside the nucleus**12.** Atomic Mass Unit**13.** Atomic Number = Protons and Electrons**15.** Quantity of protons and electrons in the nucleus of an atom**19.** The vertical rows in the periodic table |