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Thermochemistry Vocab

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|  |  |  |  R |  | 7E |  N |  T |  H |  A |  L |  P |  Y |  O |  F |  R |  E |  A |  C |  T |  I |  O |  N |  |  H |  |  L |  |  |  |
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|  |  |  |  C |  |  |  | 8H |  E |  A |  T |  | 9E |  |  | 10E |  |  |  S |  |  |  |  |  |  T |  |  H |  |  |  |
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|  |  |  |  M |  |  |  |  |  |  |  T |  |  H |  |  |  R |  |  |  V |  |  |  |  |  |  V |  |  T |  |  |  |
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|  |  |  |  S |  |  |  X |  |  |  |  A |  |  L |  |  |  Y |  |  |  T |  | 14C |  |  |  |  P |  |  F |  |  |  |
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| **Across****5.** When 1 mole of a compound is completely burnt in oxygen gas at 298k and 1 bar pressure.**7.** A change that occurs in a system when one one mole of matter is transformed by a chemical reaction under standard conditions.**8.** Energy that can be transferred due to molecular movement.**11.** Energy due to motion.**13.** The energy needed to raise the temperature of 1g of water through 1c.**15.** The SI unit of work or energy.**16.** Measure of molecular movement.**17.** The measurements of the quantity of heat exchanged.**18.** When a reaction absorbs energy.**19.** The heat required to raise the temperature of the unit of mass of a given amount.**20.** Stored energy. | **Down****1.** The transition of energy being absorbed or released during a chemical reaction or phase change**2.** Energy cannot be created or destroyed.**3.** Amount of heat needed to boil 1g of a substance at boiling point.**4.** The study of energy transfer in the form of heat.**6.** The amount of heat necessary to melt 1 mole of a substance at its melting point.**9.** A thermodynamic quantity equivalent to the heat content of a system.**10.** The ability to work.**12.** A reaction that releases heat.**14.** A measurement tool for the amount of heat involved in a chemical reaction or other process. |