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Unit 5: Chemical Reactions

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|  |  |  |  |  |  |  |  |  |  |  |  | 4S |  O |  L |  U |  B |  I |  L |  I |  T |  Y |  |  |  |  |  |  |  |  |
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|  | 8A |  |  |  |  |  |  |  | 9S |  |  |  |  |  |  L |  |  E |  |  A |  |  |  |  |  |  R |  |  O |  |  E |
|  |  C |  |  | 10C |  |  |  |  |  Y |  |  |  |  |  |  I |  |  R |  |  T |  | 11N |  |  |  |  E |  |  M |  |  D |
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| **Across****4.** the property of a solid, liquid, or gaseous chemical substance (solute) to dissolve in a solvent**13.** opposite of reduction; the substance that gives away electrons is oxidized**14.** two compounds react, and the cations and anions of the two reactants switch places, forming two new compounds or products**15.** number assigned to an element in a chemical reaction that represents the number of electrons lost/gained by an atom of that element in the compound | **Down****1.** a solid substance produced from a solution**2.** chemical reaction in which an acid and a base react with each other**3.** an element reacts with a compound and takes the place of another element in that compound**5.** a single compound breaks down into two or more elements or compounds**6.** chemical reaction where one of the products is a precipitate**7.** opposite of oxidation; reaction in which a chemical species decreases its oxidation number, by gaining electrons**8.** list of metals ranked in order of decreasing reactivity to predict displacement**9.** two or more simple substances combine to form a more complex product**10.** reaction in a hydrocarbon reacts with oxygen to form carbon dioxide and water**11.** chemical equation which lists only the species participating in the reaction**12.** tendency of a substance to undergo chemical reaction |