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| Name: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ | Date: \_\_\_\_\_\_\_\_\_\_\_\_\_\_ |

Unit 5: Chemical Reactions

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|  | E |  | 14  D | O | U | B | L | E | R | E | P | L | A | C | E | M | E | N | T | R | E | A | C | T | I | O | N |  | E |
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| **Across**  **4.** the property of a solid, liquid, or gaseous chemical substance (solute) to dissolve in a solvent  **13.** opposite of reduction; the substance that gives away electrons is oxidized  **14.** two compounds react, and the cations and anions of the two reactants switch places, forming two new compounds or products  **15.** number assigned to an element in a chemical reaction that represents the number of electrons lost/gained by an atom of that element in the compound | **Down**  **1.** a solid substance produced from a solution  **2.** chemical reaction in which an acid and a base react with each other  **3.** an element reacts with a compound and takes the place of another element in that compound  **5.** a single compound breaks down into two or more elements or compounds  **6.** chemical reaction where one of the products is a precipitate  **7.** opposite of oxidation; reaction in which a chemical species decreases its oxidation number, by gaining electrons  **8.** list of metals ranked in order of decreasing reactivity to predict displacement  **9.** two or more simple substances combine to form a more complex product  **10.** reaction in a hydrocarbon reacts with oxygen to form carbon dioxide and water  **11.** chemical equation which lists only the species participating in the reaction  **12.** tendency of a substance to undergo chemical reaction |