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chemistry

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| **Across**  **7.** the force that will increase the speed of one kilogram mass by one meter per second each second is applied  **9.** states that the total pressure of a mixture of gases is equal to the sum of the partial pressures of the component gases  **11.** the temperature -273.15 degrees celsius is given a value of zero in the Kelvin scale  **14.** a device used to measure atmospheric pressure  **16.** a common unit of pressure  **17.** there is no net loss of kinetic energy  **18.** exactly equivalent to 760 mm Hg  **19.** expresses the relationship between pressure, volume, and temperature of a fixed amount of gas  **20.** the volume of a fixed mass of a gas varies inversely with the pressure at constant temperature  **21.** states that the volume of a fixed mass of gas at constant pressure varies directly with the Kelvin temperature  **22.** pressure of each gas in a mixture | **Down**  **1.** standard conditions of exactly 1 atm pressure and 0 degrees celsius  **2.** the force per unit area on a surface  **3.** the pressure being exerted by a force of one newton acting on an area of one square meter  **4.** spontaneous mixing of the particles of two substances caused by their random motion  **5.** the idea that particles of matter are always in motion  **6.** an imaginary gas that perfectly fits all the assumptions of the kinetic-molecular theory  **8.** liquids and gases flow  **10.** a process by which gas particles under pressure pass through a tiny opening  **12.** the pressure of a fixed mass of gas at constant volume varies directly with the Kelvin temperature  **13.** simple mathematical relationships between the volume, temperature, pressure, and quantity of a gas  **15.** a gas that does not behave completely according to the assumption of the kinetic-moleculer theory |