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chemistry

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| **Across****7.** the force that will increase the speed of one kilogram mass by one meter per second each second is applied**9.** states that the total pressure of a mixture of gases is equal to the sum of the partial pressures of the component gases**11.** the temperature -273.15 degrees celsius is given a value of zero in the Kelvin scale**14.** a device used to measure atmospheric pressure**16.** a common unit of pressure**17.** there is no net loss of kinetic energy**18.** exactly equivalent to 760 mm Hg**19.** expresses the relationship between pressure, volume, and temperature of a fixed amount of gas **20.** the volume of a fixed mass of a gas varies inversely with the pressure at constant temperature**21.** states that the volume of a fixed mass of gas at constant pressure varies directly with the Kelvin temperature**22.** pressure of each gas in a mixture  | **Down****1.** standard conditions of exactly 1 atm pressure and 0 degrees celsius **2.** the force per unit area on a surface**3.** the pressure being exerted by a force of one newton acting on an area of one square meter**4.** spontaneous mixing of the particles of two substances caused by their random motion**5.** the idea that particles of matter are always in motion**6.** an imaginary gas that perfectly fits all the assumptions of the kinetic-molecular theory**8.** liquids and gases flow**10.** a process by which gas particles under pressure pass through a tiny opening**12.** the pressure of a fixed mass of gas at constant volume varies directly with the Kelvin temperature**13.** simple mathematical relationships between the volume, temperature, pressure, and quantity of a gas**15.** a gas that does not behave completely according to the assumption of the kinetic-moleculer theory |