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stoichiometry

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| **Across**  **3.** formula using the simplest whole-number ratio  **4.** reactant that is not used up when a reaction is run to completion  **7.** a number on the same line of text. In chemistry is the number to the left of a substance in a chemical equation.  **8.** shows which elements are in the compound and the actual number of each  **9.** actual yeild of a product as a percentage  **12.** the number of grams in one mole of substance. It is calculated by adding up the atomic masses of all elements in the compound. The units are always g/1 mol.  **13.** compares the actual and theoretical yield (actual/theoretical)x100; gives you an idea about the efficiency of a reaction.  **14.** a chemical reaction that has the same number of atoms of each element on both sides of it; even if they are arranged differently on each side.  **15.** the substances being made in chemical reaction; the substances on the right side of the arrow in a chemical equation. | **Down**  **1.** used up when reaction is run to completion  **2.** ratio of moles of one substance to another  **5.** maximum amount of a product that can be given off  **6.** process of using a alanced equation to determine the realitive masses reactants and products involved in the reaction  **10.** the amount of a pure substance that contains the same number of units  **11.** the ingredients in a chemical reaction; the substances on the left side of the arrow in a chemical equation. |